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***In a solution, there is 111.0 mL (110.605 g) solvent and 5.24 mL (6.0508 g) solute present in a solution. Find the mass percent, volume percent and mass/volume percent of the solute. With the solution shown in the picture below, find the mole percent of substance C. A 1.5L solution is composed of 0.25g NaCl dissolved in water. Find its molarity.***

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***Used in determining how colligative properties vary with solute concentration; masses and molecular mass of solute are known. mass percentage (%) [mass of solute (g)/mass of solution (g)] × 100: Useful when masses are known but molecular masses are unknown. parts per thousand (ppt) [mass of solute/mass of solution] × 10<sup>3</sup> (g solute/kg solution) Used in the health sciences, ratio solutions ...***

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***Problem #19: Calculate the freezing point of a solution of 5.00 g of diphenyl C<sub>12</sub>H<sub>10</sub> and 7.50 g of naphthalene, C<sub>10</sub>H<sub>8</sub> dissolved in 200.0 g of benzene (fp = 5.5 °C) Solution. There is a tiny***

**curve in this problem, but keep in mind that colligative properties are all about how many particles in solution and nothing else.**

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**10. Van't Hoff factor ( $i$ ) is the ratio of the observed value of the colligative property in solution to the theoretically calculated value of the colligative property. (a) A non-volatile solute undergoes dissociation, then  $i > 1$ . (b) A non-volatile solute undergoes association, then  $i < 1$ .**  
**2 | Chemistry-XII. Some Important Formulae. 1. Mole fraction ( $X$ ) If the number of moles of A and B are ...**

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**Which colligative property is preferred for the molar mass determination of macro molecules ?**

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**Why ? Ans. Osmotic pressure is preferred over all other colligative properties because : (a) even in dil. soln the o. p. values are appreciably high and can be measured accurately. (b) o. p. can be measured at room temp. on the other hand elevation in B. P. is measured at high temp. where the solute may ...**

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**III. Colligative properties depend on the number of solute molecules not on the nature of the solute molecules. IV. When two solutions of different concentrations are separated by a semi-permeable membrane, osmosis allows the only solvent molecules to pass through the membrane. e. I only f. II only. 62 g. I and II h. None. 10. Barium has a body ...**

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**10. In an experiment it showed that 10 mL of 0.05**

***M solution of chloride required 10 mL of 0.1 M solution of AgNO<sub>3</sub>, which of the following will be the formula of the chloride (X stands for the symbol of the element other than chlorine) (a) X<sub>2</sub>Cl<sub>2</sub> (b) XCl<sub>2</sub> (c) XCl<sub>4</sub> (d) X<sub>2</sub>Cl (Karnataka NEET 2013) 5.***

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***The solution contained 3.50mg of protein dissolved in sufficient water to form 5ml solution. The osmotic pressure of the solution at 25C was found to be 1.54 torr. Calculate the molar mass of the protein. a. 6.2 x10<sup>3</sup> g/mol b. 5.5 x10<sup>5</sup> g/mol c. 17.9x10<sup>4</sup> g/mol d. 8.5 x10<sup>3</sup> g/mol 15. Which of the following is/are intensive properties? I. Temperature II. Pressure III. Composition IV. Mass a. I only ...***

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